

Chemactivity 6 Atomic Size Answers

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CHEMACTIVITY 6 1.The number of the valence shell corresponds to the number of the row of the Periodic Table. 2.Because the nuclei gain one proton as one moves from left to right across a period but the number of inner shell electrons remains constant.

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6.IE of Kr > IE of Br because they are in the same valence shell and Kr has the higher core charge (+8 vs. +7). IE of Rb is the lowest because core charge is +1 and its valence shell (n = 5) is larger than the valence shell (n = 4) of Kr and Br.

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6.What is the significance of the atomic number, Z, above each atomic symbol in the periodic table? 7. Based on your answer to CTO 6, what do all nickel (Ni) atoms have in common? ChemActivity 1 The Nuclear Atom 3

The Nuclear Atom

Answers to ChemActivity 1 - The Nuclear Atom- Chapter 5 1. 6.2, 6.7,7 3. 6.6,7 4. (a) neutral same e and p, ion different e and p (b) assign +1 to each proton and -1 to each electron and take the difference.

Chemactivity 7 Answers

Periodic trends predict differences between elemental characteristics as you move across the periodic table. Trends are based on Coulomb's law which mathematically relates several characteristics of an elements. Atomic size measured the distance between the nucleus of an atom and the outermost non-valence electrons of the atom. Atomic size decreases from left to right, because the addition of ...

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Group trends in atomic size (6.3) As the atomic number increases within a group the charge on the nucleus increases and the number of occupied energy levels increases. Explain the group trend of atomic size (6.3) Atomic size increases ecause the shielding effect is greater than the effect of the increase in nuclear charge.

Chapter 6: The Periodic Table Flashcards | Quizlet

Atomic Structure. 1 The Nuclear Atom 2. 2 Atomic Number and Atomic Mass 8. 3 Coulombic Potential Energy 16. 4 The Shell Model (I) 22. 5 The Shell Model (II) 30. 6 Atomic Size 40. 7 Electromagnetic Radiation 44. 8 Photoelectron Spectroscopy 48. 9 The Shell Model (III) 56. 10 Electron Configurations 60. 11 Electron Configurations and the Periodic ...

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Atomic Radii Size: 1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together. 2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right. 3.

ANSWER KEY Trends on the Periodic Table

Atoms, atomic number and size are factors that define an element and depending on the atomic size or number, an element can be placed as a solid, gas or liquid in a periodic table. This is a good challenge for you if you just studied the periodic table. Give it a shot!

Lesson 10 Test - Atomic Size And The Periodic Table ...

ChemActivity 2 Atomic Number and Atomic Mass Are All of an Element's Atoms Identical?) WARM-UP Model 1: The Average Mass of a Marble In a collection of four marbles, 25% of the marbles ha ve a mass of 5.00 g and 75% of the marbles have a mass of 7.00 g. 5.00 g 7.00 g 26.00 g The average mass of a marble can be determined by dividing the total mass of the marbles by the total number of marbles ...

Solved: ChemActivity 2 Atomic Number And Atomic Mass Are A ...

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Take Krypton for example. On top it is written 36 and down bottom it is written 83.80. 36 is the atomic number and 83.80 is atomic weight or atomic mass. the easy way to remember is that the number on TOP is atomic number and bottom is the weight or mass. HOPE IT HELPED :)

Is there any site or pdf file that has the answers for ...

To the Student 1 Atomic Structure 1 The Nuclear Atom 2 2 Atomic Number and Atomic Mass 8 3 Coulombic Potential Energy 16 4 The Shell Model (I) 22 5 The Shell Model (II) 30 6 Atomic Size 40 7 Electromagnetic Radiation 44 8 Photoelectron Spectroscopy 48 9 The Shell Model (III) 56 10 Electron Configurations 60 11 Electron Configurations and the Periodic Table 64 12 Electron Spin 68 Molecular ...

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